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8. If 525 mL of 0.80 M HCl solution is

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neutralized with 315 mL of  $\text{Sr}(\text{OH})_2$  solution what is the molarity of the  $\text{Sr}(\text{OH})_2$ ?  
 $\_\_ \text{HCl} + \_\_ \text{Sr}(\text{OH})_2 \rightarrow \_\_ \text{SrCl}_2 + \_\_ \text{H}_2\text{O}$   
Solution Stoichiometry Name \_\_\_\_\_  
Chem Worksheet 15-6 example How many grams of solid calcium hydroxide,  $\text{Ca}(\text{OH})_2$ , are required to react with 350 mL of 0.40 M HCl?

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produced from the reaction of 2.50 L of  
0.250 M Calcium chloride with and  
excess of phosphoric acid?

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1] How many grams of calcium phosphate can be produced from the reaction of 2.50 L of 0.250 M Calcium chloride with an excess of phosphoric acid?

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Stoichiometry problems: 1. How many  
grams of silver chromate will precipitate  
when 150. mL of 0.500 M silver nitrate  
are added to 100. mL of 0.400 M



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potassium chromate?  $2 \text{AgNO}_3(\text{aq}) + \text{K}_2\text{CrO}_4(\text{aq}) \rightarrow \text{Ag}_2\text{CrO}_4(\text{s}) + 2 \text{KNO}_3(\text{aq})$  ...

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when 150. mL of 0.500 M silver nitrate are added to 100. mL of 0.400 M potassium chromate?  $2 \text{AgNO}_3(\text{aq}) + \text{K}_2\text{CrO}_4(\text{aq}) \rightarrow \text{Ag}_2\text{CrO}_4(\text{s}) + 2 \text{KNO}_3(\text{aq})$

## **Solution Stoichiometry Worksheet**

Stoichiometry Involving Solutions

Worksheet. 1. Calculate the number of

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mL of 2.00 M  $\text{HNO}_3$  solution required to react with 216 grams of Ag according to the equation. ... the minimum volume of the  $\text{Na}_2\text{SO}_4$  solution needed to precipitate the  $\text{Ba}^{2+}$  ions from the  $\text{BaCl}_2$  solution. 6.

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Sections 3.9 & 10: Solution

Stoichiometry: You must consider the balanced chemical reaction and set up the appropriate conversion factors to solve solution stoichiometry and titration problems. Do not use the dilution formula for these problems! Example. How many L of 0.164 M  $\text{Ca}(\text{OH})_2$  is needed to neutralize 25.00 mL of 0.458 M HCl

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Directions: Solve the following  
hypothetical stoichiometry problems.  
Assume water is available. 1. When 6.5

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mol of solid potassium chlorate breaks into solid potassium chloride and oxygen gas, what mass of solid potassium chloride is produced? 2.

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Calculate the molarity of the  $\text{H}_2\text{SO}_4$   
solution if it takes 40.0 mL of  $\text{H}_2\text{SO}_4$  to  
neutralize 0.364 g of  $\text{Na}_2\text{CO}_3$ . 0.0859

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