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Percent Yield 9. Liquid nitroglycerine (C₃H₅(NO₃)₃) is a powerful explosive. When it detonates, it produces a gaseous mixture of nitrogen, water, carbon dioxide, and oxygen. What is the theoretical yield of nitrogen 5.55 g of nitroglycerine explodes? If the actual amount of nitrogen obtained is 0.991 g, what is the percent yield of nitrogen?

Limiting Reagents - Ms. Mogck's Classroom - Ms. Mogck's ...

Holt ChemFile: Problem-Solving Workbook 48 Mole Concept Mole Concept Suppose you want to carry out a reaction that requires combining one atom of ... Yes; the answer has the correct units of atoms of boron. Is the number of significant figures correct? Yes; the mass of boron was given to three significant figures. 1 mol B

Skills Worksheet Problem Solving

CHEMFILE MINI-GUIDE TO PROBLEM SOLVING CHAPTER 10 ... •Is the answer reasonable? Yes; 83% is about 5/6, which ap-pears to be close to the ratio 0.71/0.86. 1. Calculate the percent yield in each of the following cases: ... (Hint: Percent yield can also be calculated by dividing the actual yield in moles by the theoretical yield in moles.)

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