

## Determination Of Surface Pka Values Of Surface Confined

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### Determination Of Surface Pka Values

a theoretical value for the surface PKa about 1 pH unit higher than the solution pH at the capacitance maximum (pH 3.6) for the conditions used to obtain the data in Figure 2. Therefore, we estimate that the 4-MP surface PKn is  $4.6 \pm 0.5$  at +0.2 V. This value can be compared to the bulk-phase PKa of 4-MP which is 1.4.20 This surprisingly

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Assembled Monolayers on Silicon Surfaces. Langmuir 2012, 28 (7) , 3403-3411. DOI: 10.1021/la203560k. Ting Chen, Pei-Xia Dai, Jing-Yi Wu, Dong Wang, and Li-Jun Wan .

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### **Determination of polyelectrolyte pKa values using surface ...**

The determination of pKa values of acids can be done in a multitude of manners, ranging from the classic potentiometric titration to computationally. We envisioned that it would be possible to determine the pKa values of polyelectrolytes using surface tension measurements at the surface-to-air interface. When measuring surface tension at the surface-to-air interface of polyelectrolytes at differing solution pH values a classic sigmoidal curve was obtained with the inflection point being ...

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$K_{eq} = 10^{pKa(\text{product}) - pKa(\text{reagent})}$  If you remember that your  $K_{eq}$  is always going to be 10 to the power of the right pKa value minus the left pKa value, you'll always be able to easily estimate the equilibrium constant and tell if your reaction is favorable or not.

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Determining pK a values from titration curves or free energy calculations. From the titration of protonatable group, one can read the so-called pK a 1/2 which is equal to the pH value where the group is half-protonated.

### **Protein pKa calculations - Wikipedia**

Determination of Surface PKa Values of Surface-Confined ... Determination of Surface pKa by the Combination of Neutron Reflection and Surface Tension Measurements. Langmuir 1997 , 13 (26) , 6881-6883. DOI: 10.1021/la970906r. Determination of surface pKa values of surface-confined ...

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Titration curves in aqueous medium and I ¼ 0.15 M and 25 C of (a) methylparaben, pK 00 a ¼ 8:15 at I ¼ 0.15 M and 25 C (c t ¼ 0.062 M); (b) piperazine dihydrochloride, pK 00 a ¼ 5:66 and pK ...

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Determination of pKa value by Half Neutralization/ Henderson Hasselbalch equation The mixture acts as a buffer solution. During the course of titration acid concentration decreases and salt concentration goes on increasing. During any stage of titration the pH of buffer solution is explained by the Henderson -Hasselbalch equation.

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