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Section 3 Predicting The Products Of Chemical Reactions

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High School. Section 3: Predicting the products of chemical
reactions Predict the products of the following reactions: $_ \text{A9} +$
 $_ \text{CuSO}_4$

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$2 \text{KClO}_3 (\text{s}) \rightarrow 3 \text{O}_2 (\text{g}) + 2 \text{KCl} (\text{s})$ The potential products in

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double-replacement reactions are simple to predict; the anions and cations simply exchange. Remember, however, that one of the products must precipitate, otherwise no chemical reaction has occurred.

5.6: Predicting Products from Chemical Reactions ...

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Look at the reactants of a chemical reaction and predict the products. If you got the correct compounds but in a different order than was the correct answer then use the "I was RIGHT" button. 3 things to always remember when balancing chemical equations: 1) $HX + Y(OH) \dots$

Unit 1.3: Predicting Products of Chemical Reactions ...

Also, this is an acid-base reaction, so the products should be salt and water. 11. Decomposition $Al_2(CO_3)_3 + \text{heat} \rightarrow Al_2O_3 + 3CO_2$. Another decomposition that will generate CO_2 and a metal oxide (note that Al forms a +3 ion, Al^{3+} , and monatomic oxygen will have a -2 charge, O^{2-}). 12. Synthesis $4Al + 3O_2 \rightarrow 2Al_2O_3$

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Section 3 Predicting The Products Of Chemical Reactions

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3 Predict if a reaction will occur when you combine aqueous solutions of iron (II) chloride with aqueous sodium carbonate

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solution. If the reaction does occur, write a balanced chemical equation showing it. (aq) (aq) PREDICTING REACTION PRODUCTS: DOUBLE REPLACEMENT REACTIONS Using a SOLUBILITY TABLE: Both reactants are soluble SO REACTION OCCURS!

Steps to Predicting the Products of Chemical Reactions

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The first is predicting the formation of a theoretically impossible product such as NaCO_3 or Ag_4Cl . The second is failing to balance the equation once the products have been accurately predicted Here are some tips you may find handy in helping to predict the type of reaction that will occur if you know only the reactants.

Chemistry: Predicting Reaction Products

Answers for Predicting Products of Chemical Reactions For all combustion reactions of hydrocarbons or hydrogen-carbon-oxygen molecules, the products will always be CO_2 and H_2O (assuming a complete reaction). a. Combustion: $\text{C}_6\text{H}_{12} + 9\text{O}_2 \rightarrow 6\text{CO}_2 + 6\text{H}_2\text{O}$ b. Combustion: $2\text{C}_4\text{H}_6 + 11\text{O}_2 \rightarrow 8\text{CO}_2 + 6\text{H}_2\text{O}$ c. Combustion: $\text{C}_6\text{H}_{10}\text{O}_3 + 7\text{O}_2 \rightarrow 6\text{CO}_2 + \dots$

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Section 2: Predicting the products of chemical reactions • Predict the products of the following reactions. • Predict the states of matter for double replacement reactions. • Finally balance each reaction. 1) $\text{Ag} + \text{CuSO}_4 \rightarrow$ 2) $\text{NaI} + \text{CaCl}_2 \rightarrow$ 3) $\text{O}_2 + \text{H}_2 \rightarrow$ 4) $\text{HCN} + \text{CuSO}_4 \rightarrow$ 5) $\text{HNO}_3 + \text{Fe}(\text{OH})_3 \rightarrow$ 6) $\text{LiBr} + \text{Co}(\text{SO}_3)_2(\text{aq}) \rightarrow$ 7) $\text{LiNO}_3 + \text{Ag} \rightarrow$ 8) $\text{H}_2\text{CO}_3 \rightarrow$ 9) $\text{H}_2\text{SO}_4 + \text{NH}_4\text{OH} \rightarrow$ 10) $\text{C}_4\text{H}_8 \dots$

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